



SYNTHESIS OF SCHIFF BASE METAL COMPLEXES: A CONCISE REVIEW

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Keywords: Schiff base, metal ion, dentation, synthesis, coordination.

Schiff bases are a group of compounds prepared by the condensation of primary amines and active carbonyl compounds. The Schiff bases have a general structure $RR'C=N-R''$ (where R, R' and R'' are alkyl, cyclohexyl, hydroxyalkyl, hydroxyaryl, etc). Herein the different synthetic routes of Schiff Base complexes like, direct synthesis, in situ method, oxidation of coordinated secondary amine, amine exchange approach, metal exchange and ligand exchange reaction are reviewed. A concise survey of literature on the coordination modes of complexes of Schiff base ligands is also presented in this article. The coordination complexes have been successfully screened against different strains of bacteria where they depicted the potential antimicrobial behaviour.

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Many metallo-enzymes are involved in diverse metabolic pathways including DNA synthesis, sugar metabolism and protein modification.²

Coordination chemistry is now recognized as an independent discipline covering a wide range of areas from medicine to environment. In recent past there has been a great upsurge in the studies of metal complexes of bioinorganic medicinal relevance. The discovery and basic concepts of medicinal inorganic chemistry have recently reviewed.³⁻⁵ The field now encompasses active metal complex, metal ions and even metal ions can be removed from a biological system by judicious use of metal binding molecules (termed ligands from the Latin word *ligare*, meaning that which binds).

Medicinal inorganic chemistry is at the interface between medicine and inorganic chemistry, and includes metal-based drugs, metal sequestering and mobilizing agents and metal containing diagnostic aids.^{6,7} In the early systematic study of metals in medicine (during the early to middle part of the twentieth century), recognition of the essentiality of some metal ions (e.g. iron, zinc and copper) for the avoidance of deficiency disease was a major step forward. Not only some many metal ions are essential nutrients, but many are also becoming increasingly prevalent components of diagnostic or the therapeutic agents to study or treat wide variety of disease and metabolic disorders.^{3,7,8} The list of metal ions that qualify for essential status is a work in progress; it includes not only expected members such as zinc, copper and manganese but also man formerly thought of only a poisonous, such as selenium and molybdenum.^{9,10} Included in the "possibly essential" list are such unexpected candidates as arsenic, nickel, silicon, and vanadium.¹⁰⁻¹²

The applications of metal complex in qualitative and quantitative chemical analysis have also been the subject of numerous studies during the past decade.¹³ To understand the level of pollution, analysis or measurement of the level of pollution is undertaken called the environment pollution analysis. This indicates the exact degree of pollution. Coordination compounds have found a very important place in environmental chemistry for determination of pollutants by complexometric and spectrophotometric methods.

Introduction

The chemistry of coordination compounds is well known and widely studied, yet many new developments are continuously taking place in this area of inorganic chemistry. The preparative work of Jorgenson, Werner and many others during the turn of the century opened up new vistas in coordination chemistry of transition metals. It was Alfred Werner¹ who systematized the subject coordination chemistry, by propounding theory in 1893, for which he awarded Nobel Prize in 1913. These ideas have considerably enriched the understanding of the nature of metal-ligand bonds, structure and chemistry of metal complexes, their stabilities and liabilities and other properties. It is now the need of time that new ligands of specified design should be synthesized which could lead to metal complexes with special desired and possible predictable properties. The steady improvement in synthetic methodology allows us to foresee coordinating chemistry entering a phase of creative chemistry.

For a long time the coordination compounds were considered as a rare and special class but now they have been recognized as the most resourceful type of compounds which help the nature to convert simple inorganic molecules into organic matter. For instance CO_2 and H_2O are converted into sugars by coordination compound, called chlorophyll, the magnesium porphyrin complex present in plants. A similar iron porphyrin complex present in hemoglobin of animals operates in the red blood cells as a carrier of oxygen for the oxidation of organic matter to produce carbon dioxide water and energy. Hemoglobin is also concerned with the transport of CO_2 and acid base balance of the body.

The development of complexes for application in medicine is an obvious example of investigation and creativity. Remarkable example is *cis*-platin, *cis*-PtCl₂(NH₃)₂ introduced by Rosenberg,¹⁴ which is successfully employed world wide as anticancer drugs.^{15,16} Use has been made of the calcium complexed EDTA as a scavenger for removing heavy metals like Pb(II), Cd(II) etc in case of such heavy metal poisoning in human. Antitumour activity of Cu(I) and Cu(II) complexes have attracted much attention.¹⁷ Another important use of metal complexes and chelating molecules is in the field of therapy. The use of gold complexes in the treatment of tuberculosis dates back to 1917 and many complexes of gold are known to be useful in the treatment of arthritis since 1927. In the course of study of a number of lipid soluble gold complexes for possible application of arthritis therapy, Winstock and coworkers¹⁸ prepared an unusual complex of gold. Many drugs are known to be potentiated by complexation with Mn⁺ ions. Thus, Fe(III) increases antibacterial activity of 8-hydroxyquinoline significantly, antitubercular activity of isonicotinic acid hydrazide is increased tenfold by Cu(II), carcinostatic activity of many drugs is enhanced by chelation.¹⁹

The insulin-like effect of vanadium salts in cells^{20,21} and diabetic animals²²⁻²⁸ is been known since the 1980s. Diabetic patients frequently have both abnormal glucose and lipid metabolism that can be normalized by treatment with insulin. Studies testing compounds in animal model systems²³⁻²⁸ and in human beings²⁹⁻³² show that simple vanadium salts and vanadium complexes alleviate the symptoms of diabetes.

Coordination compounds involving Schiff bases

In 1864, Schiff³³ prepared a series of compounds by the condensation of primary amines and active carbonyl compounds and since then they are known as Schiff bases. The Schiff bases have a general structure RR'C=N-R'' (where R, R' and R'' are alkyl, cyclohexyl, hydroxyalkyl, hydroxyaryl, etc). The Schiff bases contain an azomethine group and have azomethine moiety (C=N). The donating properties of the lone pair of electron increases when a functional group such as -SH or -OH is sufficiently near to the azomethine moiety and this facilitates the formation of a stable metal complex.

The coordination chemistry of the metal chelates of Schiff bases is known for more than a century. In 1840, Ettling³⁴ prepared the copper(II) complex with salicylideneimine. Since then a good amount of literature³⁵⁻⁵² on Schiff bases and their metal complexes have appeared.

Lager in 1963 revived the synthesis of Schiff bases and also studies their physical and chemical properties.⁵³ Dwyer and Mellor⁵⁴ described their chelating characteristics in a monograph. Holmes and co-workers⁵⁵ have presented an excellent review of the metal complexes of Schiff bases and Bayer indicated a novel approach to the structure and specificity of the organic chelating agents including some of the Schiff bases. Jungeis *et al.*⁵⁷ in 1969 have described the detailed applications of Schiff bases in inorganic analysis.

Synthetic routes to Schiff base complexes

Major development in the chemistry of transition metal complexes of Schiff bases took place in the last two decades and much emphasis had been laid on synthesis and structure. The synthesis, properties and structure of the Schiff base complexes depend upon nature of the Schiff base and metal ion. The following methods have generally been used for the synthesis of metal complexes of Schiff base.⁵⁸

Direct synthesis

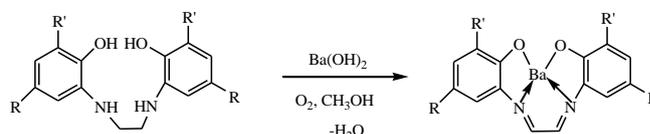
In this method, the Schiff base is allowed to react with metal ion in a suitable solvent. In order avoid the hydrolysis of azomethine group the use of organic solvent is preferred, although in a number of cases the use of binary azeotropic mixture of water and an organic solvent has been reported.

In situ method

In this method the metal ion is added during or shortly after the mixing of the aldehyde and amine. In a number of cases, it has been reported that the metal ion coordinates with one component and then reacts with other components, thus facilitating the reaction of complexation.

Oxidation of coordinated secondary amine

Here the metal ion facilitates the oxidative dehydrogenation of the secondary amine and the Schiff base complexes are obtained directly.⁵⁹ One such reaction is given in scheme 1.

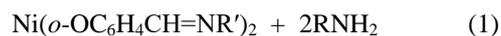


where R = H, n-C₄H₉, C₆H₅, R' = C₄H₉, C₆H₅, o-HSC₆H₄

Scheme 1. Formation of Schiff base from secondary amine

Amine exchange reaction

In this method, an amine reacts with transition metal complex of Schiff bases and exchange of alkyl/aryl groups take place as illustrated below.⁶⁰

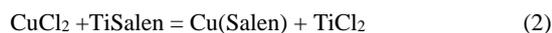


where R = H, n-C₄H₉, C₄H₁₀, C₆H₅, o-HSC₆H₄

Metal exchange reaction

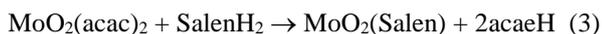
This method is used for the preparation of Schiff base complexes, which are difficult to obtain by any other method.

A typical example of metal exchange reaction is illustrated below.



Ligand exchange reaction

This method is preferably used for the preparation of dioxomolybdenum(VI) complexes that are extremely difficult to synthesize by other methods. A typical reaction is given below.⁶¹



where acacH = acetylacetone

Coordination modes of Schiff base ligands/mixed ligands

It is well known that the Schiff base/mixed ligands coordinates as bi-, tri-, tetra-, penta-, hexa- and hepta-dentate ligands. A brief survey of literature on the coordination complexes of ligand is presented below.

Coordination complexes with bidentate Schiff base ligands / mixed ligands.

The Schiff base derived from aromatic *o*-hydroxyaldehyde and primary amine behave as bidentate ligands. Generally the stereochemistry of the bidentate Schiff base complexes depend upon the group present in the Schiff base on increasing the size of the R group the stereochemistry changes from square planar to tetrahedral. The Co(II) complexes of bidentate Schiff base (Figure 1) (R = H, OH) are square planar.

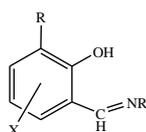


Figure 1. A bidentate Schiff base ligand.

Topich⁶² has achieved synthesis of a heterochelate complex $\text{MoO}_2(\text{Acac})\text{L}$ (where LH = Schiff base) by the reaction of $\text{MoO}_2(\text{acac})_2$ and Schiff base (Figure 2) in ethanol or ethyl acetate.⁶³ Only one acetylacetone moiety was not displaced even when the metal to ligand ratio was greater than unity. IR data indicates that the ligand coordinates through to oxygen atom of the carboxylic acid group and nitrogen atom of azomethine group and not through the phenolic oxygen atom.

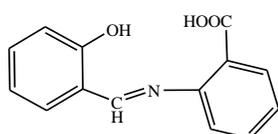


Figure 2. A Schiff's base ligand

Dioxomolybdenum(VI) complexes of the type $\text{cis-MoO}_2\text{L}_2$ have also been prepared with the Schiff bases derived from hydroxylamine and *o*-hydroxyacetophenone or 2,4-dihydroxyacetophenone.⁶⁴ Hill et al.⁶⁵ have described the synthesis of complex of the type MoO_2L_2 with the Schiff bases (Figure 3 and 4) by the reaction of bis(acetylacetonato)dioxomolybdenum (VI) and the Schiff bases in THF. The Schiff base (4) yielded a tetrameric complex of the type $\text{Mo}_4\text{O}_{11}\text{L}_2$. The ligands in complexes of the type $\text{cis-MoO}_2(\text{acac})\text{L}$ are (5) and (6). Although the ligands are potentially tridentate, they behave as OH donor bidentate monobasic and NS donor bidentate monobasic ligands in $\text{cis-MoO}_2\text{L}_2$ and only signal due to the azomethine proton in the NMR spectra of dioxomolybdenum (VI) complexes (3) and (5) confirms the presence of a *cis-MoO}_2 structure.*

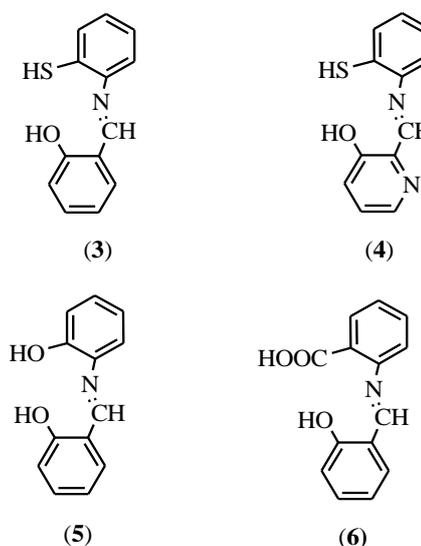


Figure 3-6. Schiff's bases forming complexes with dioxomolybdenum (VI).

Synthesis of two series of metal chelates⁶⁶ viz., (i) mononuclear chelates of composition $[\text{VO}(\text{L}_1)_2]\text{H}_2\text{O}$ and $[\text{WO}_2(\text{SCN})\text{C}_2\text{H}_5\text{OH}]_2(\text{L}_1)$ (where $\text{L}_1\text{H} = \text{N}$ -(4'-butyrylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5-one)-*p*-anisidine, 9BUMPHP-AH), N-(4-butrylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5'-one)-*m*-toluidine (BUMPHP-MTH), N-(4-butrylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5'-one)-*o*-phenylenediamine (BUMPHP-OPHDH₂), N,N'-bis(4'-butyrylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5'-one)-*m*-phenylenediamine (BMPH-MPHDH₂), or $\text{VO}(\text{L}_2')(\text{H}_2\text{O})$ (where $\text{L}_2'\text{H}_2 = \text{N}, \text{N}'$ -bis(4'-butyrylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5'-one)-*m*-phenylenediamine (BMPHP-MPHDH₂), and binuclear chelates of composition, $[[\text{VO}(\text{H}_2\text{O})(\text{OHO})_2]\text{L}_2]$ and $[[\text{WO}_2(\text{SCN})(\text{H}_2\text{O})_2]\text{L}_2]$, where $\text{L}_2\text{H}_2 = \text{N}, \text{N}'$ -bis(4'-butyrylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5'-one)-*p*-phenylenediamine (BMPH-PPHDH₂) or N,N'-bis(4'-butyrylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5'-one)benzidine (BMPHP-BZh₂) have been reported.⁶⁷

Six new binuclear dioxomolybdenum(VI) complexes or the composition $[[\text{MoO}_2(\text{O}_{11})_2]\text{L}]_2$, where $\text{L} = \text{N}$ -(4'-benzoylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5'-one)sulphamethoxazole (BMPHP-SMZ, N-(4'-benzoylidene-3'-

methyl-1'-phenyl-2'-pyrazoline-5-one)sulphadimidine (BMPHP-SIAM), N-(4'-benzoylidene-3-methyl-1-phenyl-2'-pyrazoline-5'-one)sulphadiazine (BMPHF-SDZ), N-(4'-benzoylidene-3'-methyl-1-phenyl-2'-pyrazoline-5'-one)-sulphanilamide (BMPHP-SNM), N-(4'-benzoylidene-3'-methyl-1'-phenyl-2'-pyrazolin-5'-one)sulphamerazine (BWIPT-TP-STVIR) or benzoylidenethyl-1'-phenyl-2'-pyrazolin-5'-one)sulphadimethoxine (BMPHP-SMX) have recently been reported by Maurya et al.⁶⁷ These complexes are reported to be obtained by the interaction of bis(acetylacetonato)dioxomolybdenum(VI) with the said sulpha drug based ligands in ethanol.

A new series of novel mixed-ligand complexes of Cu(II), Ni(II), Co(II), Zn(II), Sm(III) and U(VI)O₂ with the Schiff base derived from salicylaldehyde and the sulfadiazine sulfamerazine, [N-(salicylidene)sulfamerazine] (LH) (Figure 7) and the heterocyclic base, 2,2'-bipyridine (bpy) have been synthesized and characterized by Maurya et al.⁶⁸ through IR, NMR, diffuse reflectance spectra and magnetic, thermal, and molar conductance measurements. The coordination by the azomethine nitrogen is inferred by the down field shifting or the -CH=N- signal in the NMR spectra and the shift of $\nu(\text{C}=\text{N})$ to lower frequencies by 1540 cm⁻¹ in the IR spectra upon complexation. Conductance measurements confirm the non-electrolytic nature of these complexes. The presence or lattice and coordinated water molecules are indicated by thermograms of the complexes. The general compositions of the complexes were found to be [M(L)(bpy)(OAc)(H₂O)], [Sm(L)(bpy)(OAc)₂(H₂O)]₂H₂O and [UO₂(L)(bpy)(OAc)(H₂O)H₂O], where M=Cu(II), Ni(II), Co(II) or Zn(II) and HOAc=acetic acid.

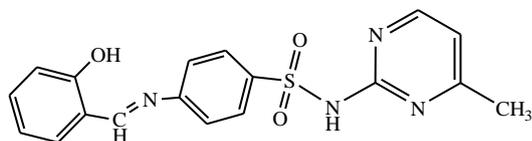


Figure 7. [N-(salicylidene)sulfamerazine] (LH).

Thompson et al.⁶⁹ have reported bis(3-hydroxy-4-pyranato)oxovanadium(IV) and tris(3-hydroxy-4-pyrone)vanadium(III) complexes, assessed them for insulin mimetic potential.⁷⁰ Key feature of these ligands are a six member ring, with a ring N or O atom either ortho or para to ketone group, and hydroxypyridinones are characterized by synthetic versatility and high affinity for a range of metal ions, rendering these ligands excellent choices for the formulation of therapeutic and/or diagnostic metallopharmaceuticals.

A series of bis(maltolato)oxovanadium(IV) (BMOV), bis(maltolato)dioxomolybdenum(VI), bis(maltolato)chromium(III), bis(maltolato)copper(II) and bis(maltolato)cobalt(II) complexes have recently reported by Thompson et al.⁷¹

In view of insulin mimetic properties of oxovanadium(IV), Mohammadi et al.⁷² have synthesized novel bis(4-hydroxy-3-methoxyphenyl)-1,6-heptadiene-3,5-dione (curcumin) complexes with the formula, ML₃, where M is Ga(III) or In(III), or of the formula, ML₂ where M is [VO]²⁺ and characterized by mass spectrometry, infrared absorption

spectroscopy and elemental analysis. The ligand curcumin behave as monobasic bidentate in these complexes.

Four new solid derivatives, [V(IV)O(1,2-diethyl-3-hydroxyl-4-pyridonate)₂], (V(IV)O(1-(p-tolyl)-2-ethyl-3-hydroxyl-4-pyridonate)₂), [V(IV)O(1-P-(n-butyl)phenyl)-2-ethyl-4-pyridonate)₂] and (V(IV)O(1-hexyl)phenyl-2-ethyl-3-hydroxyl-4-pyridonate)₂], were isolated and characterized by Garribba et al.⁷³ Aqueous solution studies regarding the identification and characterization of complexes formed by the V(IV)O ion and 3-hydroxyl-4-pyridone derivatives have been performed using EPR and UV-Vis spectroscopic techniques. Garribba et al.⁷⁴ have recently reported 2-pyrazinecarboxylic acid and three of its derivatives (5-methyl-2-pyrazinecarboxylic, 2,3-pyrazinedicarboxylic and 5-hydroxypyrazinecarboxylic acids) coordinate the V(IV)O ion forming VOL, VOL₂ and (VO)₂L₂H₂ species in acidic and neutral solutions. Bis chelated species are hexacoordinated and are characterized by a cis/trans isomerism, with the trans arrangement favoured with respect to the cis one. The ligand used in this synthesis behave as monobasic bidentate with an N,O coordination are characterized by an (N_{aromatic}, COO⁻) donor set.

Coordination complexes with tridentate Schiff base ligands/mixed ligands

The tridentate Schiff bases are obtained by the condensation of amines containing OH or SH group and o-anisidine-hydroxycarbonyl compounds or β-diketones. The ligand (Figure 8) derived from salicylaldehyde and o-aminophenol is an important example of tridentate Schiff base. The tridentate Schiff base ligands react with the metal ions to give metal complexes having novel structural, magnetic and spectroscopic properties.⁷⁵⁻⁷⁷

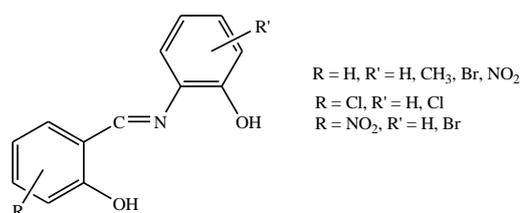


Figure 8. Substituted (E)-2-(hydroxybenzylideneamino)phenols.

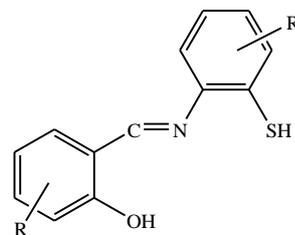


Figure 9. Schiff base forming complex with oxovanadium(IV).

The oxovanadium(IV) complexes of the type VOL (Figure 9, Substituent(s) = 5-chloro; 5-chloro, 5-bromo, 5-nitro, 3,5-dichloro, 5,6-benzo) have been reported by Lee et al.⁷⁸ The magnetic moments of these complexes are in the range 1.27-

1.32 B.M. The magnetic moments decrease considerably with lowering of temperature indicating the presence antiferromagnetic exchange. The $\nu(\text{V}=\text{O})$ stretch of the complexes lies in the range $900\text{--}1005\text{ cm}^{-1}$. The electronic spectra of the complexes exhibit a broad band in the region $13,300\text{--}20,000\text{ cm}^{-1}$ due to the d-d transition. It has been noticed that oxovanadiin(IV) complexes with ONS donor sites exhibit a greater degree of antiferromagnetism than the oxovanadium (IV) complexes with ONO donor sites.

A series of copper(II) and zinc(II) complexes involving a tridentate O,N,O-donor⁷⁹ Schiff base derived from salicylaldehyde and β -alanine [i.e. N-salicylidene- β -alanine(2-), L] having the composition $[\text{Cu}_2(\text{L}_2)(\text{H}_2\text{O})]\cdot\text{H}_2\text{O}$, $[\text{Cu}(\text{L})(\text{H}_2\text{O})]_n$ and $[\text{Zn}(\text{L})(\text{H}_2\text{O})]_n$ (**3**) have been prepared and characterized by elemental analyses. UV-Vis, FT-IR and electron spray ionization mass (ESI-MS) spectra and thermal analyses. Complexes (**1**) and (**2**) have been investigated by single X-ray analysis and also by temperature dependent magnetic susceptibility measurements (294-80 K). All complexes have been evaluated by the antiperoxynitrite activity assay and alloxan-induced diabetes model. The significant antioxidant activities have been found in the case of both Cu(II) complexes. In spite of this first two complexes, the Zn(II) complex, as well as the potassium salt of the ligand (KHL) showed only insignificant protective effect against the tyrosine nitration *in vitro*.

Coordination complexes with tetradentate bases/mixed ligands.

The tetradentate Schiff bases are obtained by the condensation of an aromatic o-hydroxycarbonyl compound and a diamine. The important examples of the quadridentate Schiff bases (Figure 10) are the condensation products of salicylaldehyde or substituted salicylaldehyde and diamines. These Schiff bases enforce a high degree of planarity to the metal chelates.⁸⁰ With the change of n from 2 to 4, the stereochemistry of the Cu(II) and Co(II) complexes of Schiff bases changes from square planar to tetrahedral.⁸¹⁻⁸³

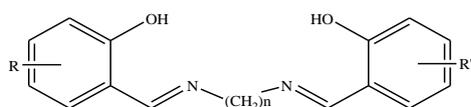
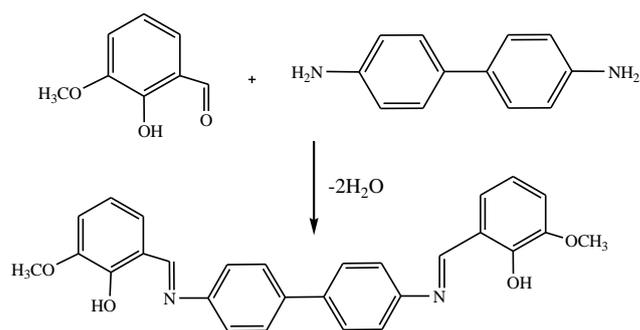


Figure 10. A tetradentate Schiff base ligand.

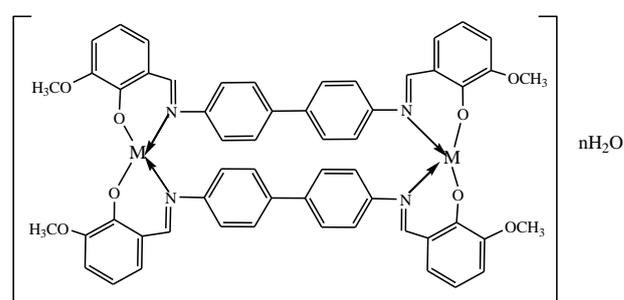
A binucleating tetradentate Schiff base ligand, bis(*o*-vanillin)benzidine (*o*-v₂bzH₂), and its seven new binuclear complexes have been synthesized (Scheme 2) and characterized on the basis of elemental analysis, IR, NMR, electronic, magnetic, thermal studies and conductance measurements. The compositions of these complexes were found to be $[\text{M}(\text{o-v}_2\text{bz})_2\cdot n\text{H}_2\text{O}]$, where $\text{M} = \text{Cu}(\text{II}), \text{Ni}(\text{II}), \text{Co}(\text{II}), \text{Zn}(\text{II}), \text{Mn}(\text{II})$ or $\text{UO}_2(\text{V1})$, and $\{\text{Sm}(\text{o-v}_2\text{bz})(\text{OAc})(\text{H}_2\text{O})\}_2$.

The ¹HNMR spectrum of one of the compounds, $[\text{Zn}(\text{o-v}_2\text{bz})_2]$ shows the absence of proton signal for phenolic oxygen (-OH). Low magnetic moment values, high thermal stability and insolubility in common organic solvents support the binuclear structure of these complexes. Suitable binuclear structures (Figure 11) have been assigned. The 3D

molecular modeling (Fig. 14) and analysis for bond lengths and bond angles have also been carried out of one of the representative compounds, $[\text{Ni}(\text{o-v}_2\text{bz})_2]$ (**2**).



Scheme 2. Synthesis of bis(*o*-vanillin)benzidine.



$\text{M} = \text{Cu}(\text{II}), n = 0, 2; \text{Co}(\text{II}), n = 0; \text{Zn}(\text{II}), n = 0; \text{Mn}(\text{II}), n = 0$

Figure 11. Structure of binuclear complexes of a tetradentate Schiff's base.

Coordination Complexes with Penta-, Hexa- and Heptadentate Schiff Base/ Mixed Ligands.

Relatively few reports on the metal complexes of the Schiff bases having potential penta-, hexa- and heptadonor sites have appeared as compared to the metal complexes with bi-, tri- and tetradentate Schiff bases. The Schiff base (Figure 12) derived from diethylenetriamine and salicylaldehyde behaves as a dibasic pentadentate (N_3O_2) donor's ligand.

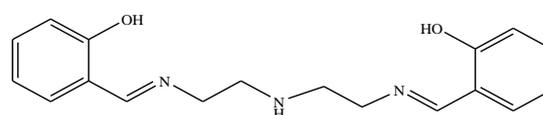


Figure 12. A dibasic pentadentate Schiff's base.

The coordination of pentadentate ligand to the central metal ion can take place only when the fifth axial donor atom present is flexible. The five-coordinated compounds can acquire two different geometries depending upon the flexibility of chelate ring and length of chain between the two coordinated atoms. Thus, the metal complex would have a trigonal bipyramidal geometry if the chelate ring is quite flexible and if the length of chain is small the resulting complex will be of square pyramidal geometry.

Dioxouranium(VI) complex of the type UO_2L (where LH_2 =Figure 12) has been prepared by the reaction of an equimolar mixture of dioxouranium(VI) acetate and Schiff base.⁸⁵ The Schiff base behaves as a neutral pentadentate ligand. UO_2X_2 and $UO_2(ClO_4)_2$ react with the Schiff base shown in Figure 12 as a result nine coordinated complexes $UO_2(LH_2)X_2$ and UO_2L (where $X=Cl^-$, I^- , NO_3^- and $0.5SO_4^{2-}$) are formed, respectively. The Schiff base behaves as a neutral pentadentate in $UO_2(LH_2)X_2$ and dibasic pentadentate ligand in UO_2L . The nitrate ion behaves as a monodentate ligand while the sulphate ion behaves as a bidentate ligand.

Dioxouranium(VI) complexes of the type UO_2L (where LH_2 =Schiff base derived from salicylaldehyde and 1,5-diamino-3-oxopentane) have been reported.⁸⁴ The complexes exist in α - and β -forms. In both the modifications uranium is pentagonal bipyramidal with the pentadentate ligand coordinated in a plane normal to UO_2 . These are conformational isomers differing mainly in the spatial arrangement of the ethylenic chains. The UO_2 group is non-linear in both the forms.

Dioxouranium(VI) complex of the type $[UO_2(LH_2)(NO_3)_2]^{2+}$ $[UO_2(NO_3)_4]$ (where LH_2 =Schiff base 2,6-diaethylpyridine bis(2'-pyridylhydrazone)) has been prepared by the reaction of uranyl nitrate and the Schiff base.⁸⁷ The Schiff base behaves as a neutral pentadentate ligand. IR data indicate the monodentate behavior of nitrate group in the cationic moiety while the bidentate behavior of two nitrate group and the remaining two as monodentate of the anionic part giving a six fold coordination of oxygen atoms in the equatorial plane normal to uranyl group.

The Schiff bases (LH_2) derived from substituted salicylaldehydes and bis(3-aminopropyl)amine or bis(3-aminopropyl)methylatninc from high-spin five coordinated complexes ML with manganese(II), cobalt(II), nickel(II) copper(II) and zinc(II).⁸⁸ The Schiff bases behave as dibasic pentadentate ligands. The nickel(II) complex when dissolved in pyridine forms hexacoordinated pyridine adducts.

Ten-coordinated dioxouranium(VI) complex, $UO_2(LH_4)X_2$ (where LH_4 = Schiff base prepared by the condensation of salicylaldehyde and triethylenetetraamine, $X=Cl^-$, I^- , NCS^- , $0.5SO_4^{2-}$) has been reported.⁸⁹ IR data indicate that the Schiff base function; as neutral hexadentate (N_4O_2 donors) ligand. The nitrate and thiocyanate groups behave as monodentate ligands while the sulphate behaves as a bidentate ligand.

A pentadentate Schiff base, 2,6,10-thiazo-1,11-bis(2-aminophenyl)undeca-1,10-diene (L) and complexes of general formula MLX_2 (where $M=Cu(II)$, $Ni(II)$, $X = Cl^-$, Br^- , I^- , NO_3^- and ClO_4^-) have been reported.⁹⁰ IR data show an interaction between halide anion of the outer coordination sphere and the complexes amino group. ESR and electronic spectral data of the copper(II) compound are consistent with a square pyramidal geometry. Since crystal ESR studies of $CuL(NO_3)_2$ and $CuLBr_2$ revealed that the copper atoms in the former compound occupy two magnetically non-equivalent places in the lattice while the copper atoms in the later compound take identical sites.

Synthesis and characterization of pentagonal bipyramidal complexes of nickel(II), cobalt(II), iron(II), manganese(II) and zinc(II) with the heptadentate base derived from 2,6-diacetylpyridine and diethylenetriamine have been

reported.⁹¹ The complexes are of the type $[ML]X_2.YH_2O$ (where $M=Co(II)$, $Ni(II)$, $Fe(II)$, $Mn(II)$, $Zn(II)$, $X=ClO_4^-$, BPh_4^- , $Y=0$ or 1). A single crystal X-ray analysis of $[CuL](ClO_4)_2$ confirms the pentagonal bipyramidal structure.

A new mononuclear copper(II) complex $[Cu(HL)(Cl)PF_6]$ having a donor⁹² ligand ($HL=N,N'$ -bis(2-pyridylinethyl)-1,3-diaminopropane-2-ol) has been synthesized and structurally characterized by X-ray crystallography. It shows a 1D chain in packing structure through intermolecular O-H...Cl strong hydrogen bonding interaction {O-H...Cl, 1.99(4) Å and the angle (O-H...Cl) 163(4)°}. The packing structure also shows π - π stacking. The morphology of sub micrometer rod of the complex has been studied by SEM.

Raman et al.⁹³ reported the synthesis, characterization and electrochemical behaviour of Cu(II), Co(II), Ni(II) and Zn(II) complexes derived from acetylacetone and *p*-anisidine and studied their antimicrobial activity. They found that neutral tetradentate N_2O_2 type complexes of Cu(II), Ni(II), Co(II) and Zn(II) synthesized using the Schiff base formed by the condensation of acetylacetone and *p*-anisidine possess square-planar geometry. All the title complexes were screened for antimicrobial activity by the well diffusion technique using DMSO as solvent. The minimum inhibitory concentration (MIC) values were calculated at 37 °C for a period of 24 h. It was found that all the complexes are antimicrobially active and show higher activity than the free ligand.

Imran et al.⁹⁴ studied the *in vitro* antibacterial studies of ciprofloxacin-imines and their complexes with Cu(II), Ni(II), Co(II) and Zn(II). These were synthesized and characterized on the basis of physical properties, conductance measurements, elemental analysis, UV-vis, infrared and nuclear magnetic resonance spectroscopy. These ligands as well as their metal complexes were also evaluated for their antibacterial activity against several bacterial strains, such as *Staphylococcus aureus*, *Bacillus subtilus*, *Salmonella typhae*, and *E. coli*. They found that metal complexes are more antibacterial as compared to uncomplexed ligands.

Nida et al.⁹⁵ studied the synthesis, characterization and anti-bacterial studies of some metal complexes of Schiff base derived from benzaldehyde and sulfonamide. The research group subjected the complexes for their antimicrobial activity against *E. Coli* and *Salmonella Typhae*. The metal complexes of Cu(II), Ni(II) and Co(II) were found to show enhanced antimicrobial activity as compared to uncomplexed ligand.

Kavitha et al.⁹⁶ reviewed biological activities of schiff base and its complexes very efficiently. The discussion involves the mechanism and importance of these compounds. This review summarizes the synthesis and biological activities of Schiff bases and its complexes, Schiff bases are versatile ligands which are synthesized from the condensation of primary amines with carbonyl groups. These compounds are very important in medicinal and pharmaceutical fields because of their wide spectrum of biological activities. Most of them show biological activities such as antibacterial, antifungal as well as antitumor activity. Transition metal complexes derived from the Schiff base ligands with biological activity have been widely studied.

Singh et al.⁹⁷ surveyed synthesis and bioactivity of some metallo-sulpha drugs as per reports of that era. They summarized that compounds containing the sulphonamide group have long been used as drugs for various diseases. The biological activity of these drugs is enhanced on undergoing complexation with metal ions.

Sahu et al.⁹⁸ also recently reviewed the summary of medicinal chemistry features of schiff base and its enhancement of biological activities on complexation.

Donde et al.⁹⁹ reported the synthesis, characterization and biological activity of mixed ligand Co(II) complexes of schiff base 2-amino-4-nitrophenol-N-salicylidene with some amino acids. The Schiff base and its mixed ligand complexes, in general, were non-hygroscopic and stable solids. The compounds were subjected to simultaneous thermogravimetric analysis to study their decomposition mechanism and thermal stability. The Schiff base and mixed ligand complexes were preliminary screened against various strains of microbes to study their biological effect.

Tharinamj et al.¹⁰⁰ brought to light the synthesis and spectral characterization of some transition metal complexes of azomethine derivative of diaminomalononitrile. New complexes of VO(II), Mn(II), Co(II), Ni(II), Cu(II), and Zn(II) were synthesized and characterized by analytical and physicochemical techniques. These metal complexes were also tested for their antibacterial and antifungal activities to assess their inhibiting potential. Metal-mediated fluorescence enhancement is observed on complexation of the azo Schiff base ligand. The synthesized compounds were investigated for nonlinear optical properties, and the surface morphology of the Cu(II) complex was studied by scanning electron microscopy.

Complexes of Co(II), Ni(II) and Zn(II) with Schiff bases derived from 4-anisaldehyde were reported by Ndahi and Nasiru 2012.¹⁰¹ The electronic spectral data reported by them indicates that the compounds are six coordinated. The molar conductance values showed that the complexes are non-electrolytes, The compounds were screened in vitro for antibacterial activity against some pathogenic bacteria: *Escherichia Coli*, *Pseudomonas aeruginosa*, *Salmonella typhi*, *Bacillus subtilis* and *Staphylococcus aureus* using the agar-well diffusion method. The synthesized Schiff base complexes exhibit higher antibacterial activity against the tested pathogens compared to the free Schiff base because of chelation.

Antibacterial studies of some Schiff base metal complexes containing Zn(II), Cu(II), Ni(II), Co(II), Mn(II), Cr(III) and Cd(II) with some novel antibiotics have recently reported by Bukhari et al.¹⁰² The antibacterial activity showed the following trend: Metal complexes > Schiff base ligands > Parent drugs.

Mahendra et al.¹⁰³ carried out synthesis, characterization and biological activities of 5-chloroisatin Schiff base and its metal complexes comprising of copper(II), cobalt(II), nickel(II) and zinc(II). The ligand and its metal complexes have been screened for their antibacterial activity against *Staphylococcus aureus*, *Escherichia coli*, etc. and antifungal activity against *Aspergillus niger*, *Aspergillus flavous*, etc.

The activities of both the samples have shown significant and noticeable changes on complexation.

Transition metal complexes with mixed nitrogen-sulphur donor macrocyclic schiff base ligand of their synthesis, spectral, electrochemical and antimicrobial studies were recently updated by Shiekh et al.¹⁰⁴ with Cu(II), Co(II), Ni(II) and Mn(II). The results indicate that the complexes are having potential antibacterial and antifungal properties.

Besides experimental studies in the above systems computational aspects are very well studied. As per the application assets of the field various measures are guessed using theoretical field. Recently Ahmad et al.¹⁰⁵ reported inhibition effects of a synthesized novel 4-aminoantipyrine derivative on the corrosion of mild steel in hydrochloric acid solution together with quantum chemical studies, electronic properties such as highest occupied molecular orbital energy (HOMO), lowest unoccupied molecular orbital energy (LUMO) and dipole moment (μ) were calculated and discussed. The results showed that the corrosion inhibition efficiency increased with an increase in the HOMO values but with a decrease in the LUMO value.

Sulfa drug based complexes have been recently also studied with so many transition metals. Sharma et al.¹⁰⁶ brought to light the coordination chemistry of Ca(II) and Mn(II) with Schiff base of sulfaguanidine [4-amino-N-{amino(imino)methyl}benzenesulfonamide' and salicylaldehyde. The ligand behaves as a bidentate with N,O donor atoms. Complexes have been characterized by elemental analysis, UV-visible and IR spectral studies.

Conclusion

A concise survey of literature on the coordination modes of complexes of Schiff base ligands is also presented in this review article. It is clear from this review article that the Schiff bases coordinate as bi-, tri-, tetra-, penta-, hexa- and heptadentate ligands to the metal ions. The five-coordinated compounds can acquire two different geometries depending upon the flexibility of chelate ring and length of chain between the two coordinated atoms. The coordination complexes have been successfully screened against different strains of bacteria where they depicted the potential antimicrobial behavior.

Acknowledgements

Authors thank the Head, Department of Chemistry, Govt. Motilal Vigyan Mahavidyalaya Bhopal for providing necessary research facilities. Facilities provided by UGC for their generous research support are also gratefully acknowledged.

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Received: 29.10.2017
Accepted: 27.11.2017.